

Name _____ Date _____ Period _____

IONIC BONDING PRACTICE

1. An ionic bond forms between a _____ and a _____.
2. Metals typically form _____, or positively charged atoms. This means that the atom _____ electrons.
3. Nonmetals typically form an _____, or a negatively charged atom. This means that the atom _____ electrons.
4. Fill in the table below:

Group Number	Group Name	# Valence Electrons	Will it GAIN or LOSE electrons to fill its outer shell?	Charge of the Ion
1	Alkali Metals			
2				
3				
4				
5				
6				
7				
8				

5. Draw the Lewis Dot Diagram for the following atoms and write what charge its ion is most likely to have in the space below the element symbol.

a. K

b. Ca

c. Al

d. Si

e. N

f. O

g. Br

e. Ar

6. If ionic bonding is an attraction between opposite charges, A) which elements will likely bond together? B) Which is already stable? C) Which could form either an anion or cation?

A.

B.

C.

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Ionic Compounds Practice

For each of the following, draw the Lewis structures, arrows to show the movement of electrons, and then write the chemical formula for the correct ionic compound. Check your work by doing the criss-cross method.

1. Calcium + Fluorine

Formula _____

2. Aluminum + Chlorine

Formula _____

3. Barium + Oxygen

Formula _____

4. Sodium + Nitrogen

Formula _____

5. Magnesium + Nitrogen

Formula _____